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~~Find Total~~

~~possible number~~

~~of isomers for~~

~~Alkanes~~

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The Equilibrium
Constant

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~~K_w pH pOH pK_a~~

~~pK_b H⁺ OH⁻~~

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halogens and

their compounds.

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and nucleophilic

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reactions.

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alkenes and
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calculations,
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Haber cycle, and
enthalpy changes
in solution.

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atoms, molecules

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equations, gas

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calculations,

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mass, solutions,

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problems caused
by nitrogen
compounds and

nitrate

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sulfur and

oxides, sulfuric
acid and

properties, and

uses of sulfuric
acid. Practice

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acids, peptides
and proteins.

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aluminum oxide,
balancing
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balancing
equations:
reactions with
oxygen, bonding

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3 oxides,

chemical

properties of

chlorine,

chemical

properties of

oxygen, chemical

properties

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chemistry

periodic table,

chemistry:

oxides,

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chlorides of
period 3
elements,
electrical

conductivity in
period 3 oxides,
electronegativit
y of period 3
oxides, ionic
bonds, molecular
structures of
period 3 oxides,
oxidation number
of oxides,

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oxidation
numbers, oxides
and hydroxides
of period 3

elements, oxides
of period 3

elements, period
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lectronegativity

, physical
properties

periodicity,

reaction of

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sodium and
magnesium with
water, and
relative melting
point of period
3 oxides.

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of

polymerization,

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and rate of

reaction.

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transfer.

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ing/extraction
methods in
different
fields. The use

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of leaching test methods is increasing in various areas including: waste treatment and disposal; incineration of waste; soil clean-up and reuse of cleaned soil; sludge treatment. This has led to (and

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may increasingly lead to) the development of a large number of very similar tests in these different fields. However, these developments are taking place with no clear understanding of their mutual

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relationships.

In view of these developments, efforts are

needed to

harmonize the

leaching

procedures that

could be adapted

for different

matrices, as

well as validate

the use of

existing tests

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in other fields.
The development
of a wide
variety of leach
ing/extraction
tests for
different
matrices is
undesirable from
a regulatory
point of view
and undesirable
for industry.
Clarity in

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testing is
crucial in produ
cer-consumer
relations. This

collective

document will

assist in

improving the

understanding of

leaching from a

variety of

sources and

will, where

appropriate,

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help to bring
together the
approaches used
in different
technical fields
and in different
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gases, liquids
and solids

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Chemical Bonding

MCQs Worksheet

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spectrum, atomic
absorption
spectrum, atomic
emission
spectrum,
molecules,
azimuthal
quantum number,

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model defects,
charge to mass
ratio of
electron,
discovery of
electron,
discovery of
neutron,
discovery of
proton, dual
nature of
matter, electron

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charge, electron
distribution,

electron radius

and energy

derivation,

electron

velocity,

electronic

configuration of

elements, energy

of revolving

electron,

fundamental

particles,

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uncertainty

principle,

hydrogen

spectrum,

magnetic quantum

number, mass of

electron,

metallic

crystals

properties,

Moseley law,

neutron

properties,

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orbital concept,
photons wave
number, Planck's
quantum theory,
properties of
cathode rays,
properties of
positive rays,
quantum numbers,
quantum theory,
Rutherford model
of atom, shapes
of orbitals,
spin quantum

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spectrum, x
rays, and atomic
number. Practice
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atomic mass,
atoms,
molecules,
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combustion
analysis,

empirical

formula,

isotopes, mass

spectrometer,

molar volume,

molecular ions,

moles, positive

and negative

ions, relative

abundance,

spectrometer,

and

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chemical

combinations,

atomic radii,

atomic radius

periodic table,

atomic, ionic

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and covalent
radii, atoms and
molecules, bond
formation,

covalent radius,
electron

affinity, electr
onegativity, ele
ctronegativity

periodic table,
higher

ionization

energies, ionic
radius,

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ionization

energies,

ionization

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table, Lewis

concept, and

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Experimental
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chromatography,
crystallization,

filter paper

filtration,

filtration

crucibles,

solvent

extraction, and

sublimation.

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questions: Gas

laws, gas

properties,

kinetic

molecular theory

of gases, ideal

gas constant,

ideal gas

density,

liquefaction of

gases, absolute

zero derivation,

applications of

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Avogadro's law,

Boyle's law,

Charles law,

Dalton's law,

diffusion and

effusion,

Graham's law of

diffusion,

ideality

deviations,

kinetic

interpretation

of temperature,

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liquids
properties, non-
ideal behavior
of gases,

partial pressure
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properties,

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matter,

thermometry

scales, and van

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of solids,
comparison in
solids, covalent

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crystalline

solids, Avogadro
number

determination,

boiling point,

external

pressure,

boiling points,

crystal lattice,

crystals and

classification,

cubic close

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Equilibrium

packing, diamond
structure,
dipole-dipole
forces, dipole
induced dipole
forces, dynamic
equilibrium,
energy changes,
intermolecular
attractions,
hexagonal close
packing,
hydrogen
bonding,

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intermolecular
forces, London
dispersion
forces, metallic
crystals
properties,
metallic solids,
metal's
structure,
molecular
solids, phase
changes
energies,
properties of

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covalent
crystals, solid
iodine

Equilibrium
structure, unit
cell, and vapor
pressure.

Instructional
Design for
Teachers, Second
Edition focuses
on the
instructional
design (ID)

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specifically for
K-12 teachers.

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introduced a

new, common-

sense model of

instructional

design to take

K-12 teachers

through the ID

process step by

step, with a

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Chapter emphasis
on preparing,
motivating, and
encouraging new
and ongoing use
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principles. This
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includes new
material on
design in
gaming,
cybercharters,
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classrooms, and
flipped
classrooms, as
well as special
considerations
for the Common
Core. Each
chapter contains
framing
questions,
common errors,
easy-to-use
rules of thumb,
clearly stated

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outcomes, and
examples showing
ID in action.

The basic model
and its
application
within
constructivism
and user-design
will help
teachers adapt
from a
behavioral
approach to a

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more open,
student-centered
design approach.

Combining basics

with strategies

to implement

this model in

the most

advanced

instructional

approaches, this

book empowers

teachers and

learners to use

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instructional

design with the

most recent

research-based

approaches to

learning.

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how ID

principles can

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moments in
positive and
practical ways.

The book can be

used for basic

ID courses and

introductory

curriculum

courses, and is

accessible to in-

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