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of a substance is 6.02×10^{23} representative particles of that substance and is the SI unit for measuring the amount of a substance. Avogadro's Number. the number of representative particles in a mole, 6.02×10^{23} . Representative Particle.

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number of representative particles in a mole, 6.02×10^{23} . refers to the species present in a substance: usually atoms, m.... the mass of one mole of a pure substance.

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Avogadro ' s Number One important property of a mole is that it means a definite number of particles just like a

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dozen means a number of particles.
While a dozen is only 12 particles a
mole is a larger number – 6.02×10^{23}
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MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287–289) 1.

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SECTION 10.1 THE MOLE: A
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CHAPTER 10: Chemical Quantities

BASICS:

- The basic unit that is used to determine the amount of a chemical substance is called a mole
- A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance
- The mole was founded by a scientist named

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BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided

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