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The plot of  $V = V(P, T)$  for a gas is shown in Fig. 1.1. Determine the expressions of the two second derivatives of the volume of this plot. (note: the principle curvatures of the surface are proportional to these second derivatives).

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Heat Capacities The generic heat capacity c  
105 bT  $C_p = a + \frac{b}{T} + \frac{c}{T^2} + \frac{d}{T^3} + \frac{e}{T^4}$  ; T2 103 The  
heat capacities of various elements and  
compounds are  $C_{p, Ags} = C_p \hat{e}$ . 8a Ø 21.30,  
b Ø 8.54, c Ø 1.51 ...

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One mole of a monatomic ideal gas at standard temperature and pressure (STP) undergoes the following three processes:

1. at constant Pressure the Temperature is doubled.
- 2.

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The plot of  $V = V(P, T)$  for a gas is shown in Fig. 1.1. Determine the expressions of the two second derivatives of the volume of this plot.

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Thus, the thermodynamic state of a system is uniquely fixed when a small number of macroscopic, independent variables are fixed. For example, consider a gas or a liquid of constant composition such as a pure gas or liquid. The three key variables are pressure,  $P$ , temperature,  $T$ , and volume,  $V$ .

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Fig 11

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Manual thermodynamics has much wider applicability. In material science, one is normally not that interested in heat and work, but interested more the state of matter and how things might change when mixed, heated, pressurized, etc.

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