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counting the s-electrons with the p-electrons as outer electrons).

Misunderstanding electron configuration for ions.

Confusing the terms 'orbital' and 'energy level'.

Dot-and-cross diagrams

Knowing the general rule for individual atoms.

Simple ionic compounds e.g.

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the two atoms bonded together and the shared pair of electrons. This example clearly shows the refinement of the GCSE ideas required at AS level. Clearly stating that the bond is the electrostatic attraction between nuclei and a shared pair of electrons, this example scored 2. 4 GCE Chemistry 8CH0 01

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Principles of Chemistry

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Question number	Answer	Mark
4	The only correct answer is D (3p subshell 6, third quantum shell 18)	A is incorrect because 2 is the number of electrons in a 3p orbital and the 3d electrons have been omitted from the third quantum shell B is incorrect because 2 is the number of electrons in a 3p orbital C is incorrect

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because the 3d electrons have been omitted from the third quantum shell

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